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In this ninth edition, the authors continue to refine their presentation of physical chemistry. The coverage of introductory topics is streamlined, and the addition of a new fundamentals chapter provides students with an overview of key principles on which the subject is based.

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Physical Chemistry by Robert A. Alberty, Robert J. Silbey, Farrington Daniels, 2001, John Wiley & Sons Inc edition, in English - 3rd ed.

Ever since Physical Chemistry was first published in 1913 (then titled Outlines of Theoretical Chemistry, by Frederick Getman), it has remained a highly effective and relevant learning tool thanks to the efforts of physical chemists from all over the world. Each new edition has benefited from their suggestions and expert advice. The result of this remarkable tradition is now in your hands. Now revised and updated, this Fourth Edition of Physical Chemistry by Silbey, Alberty, and Bawendi continues to present exceptionally clear explanations of concepts and methods. The basic theory of chemistry is presented from the viewpoint of academic physical chemists, but detailed discussions of practical applications are integrated throughout. The problems in the book also skillfully blend theory and applications. Highlights of the Fourth Edition: A total of 170 computer problems appropriate for MATHEMATICATM, MATHCADTM, MATLABM, or MAPLETM. Increased emphasis on the thermodynamics and kinetics of biochemical reactions, including the denaturation of proteins and nucleic acids. Expanded coverage of the uses of statistical mechanics, nuclear magnetic relaxation, nanoscience, and oscillating chemical reactions. Many new tables and figures throughout the text.

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Mathematics for Physical Chemistry, Third Edition, is the ideal text for students and physical chemists who want to sharpen their mathematics skills. It can help prepare the reader for an undergraduate course, serve as a supplementary text for use during a course, or serve as a reference for graduate students and practicing chemists. The text concentrates on applications instead of theory, and, although the emphasis is on physical chemistry, it can also be useful in general chemistry courses. The Third Edition includes new exercises in each chapter that provide practice in a technique immediately after discussion or example and encourage self-study. The first ten chapters are constructed around a sequence of mathematical topics, with a gradual progression into more advanced material. The final chapter discusses mathematical topics needed in the analysis of experimental data. Numerous examples and problems interspersed throughout the presentations Each extensive chapter contains a preview, objectives, and summary Includes topics not found in similar books, such as a review of general algebra and an introduction to group theory Provides chemistry specific instruction without the distraction of abstract concepts or theoretical issues in pure mathematics

Thermodynamics is the science that describes the behavior ofmatter at the macroscopic scale, and how this arises fromindividual molecules. As such, it is a subject of profoundpractical and fundamental importance to many science andengineering fields. Despite extremely varied applications rangingfrom nanomotors to cosmology, the core concepts of thermodynamicsuch as equilibrium and entropy are the same across alldisciplines. A Conceptual Guide to Thermodynamics serves as a concise,conceptual and practical supplement to the major thermodynamicstextbooks used in various fields. Presenting clear explanations ofthe core concepts, the book aims to improve fundamentalunderstanding of the material, as well as homework and exampereformance. Distinctive features include: Terminology and Notation Key: A universaltranslator that addresses the myriad of conventions, terminologies, and notations found across the major thermodynamics texts. Content Maps: Specific references to eachmajor thermodynamic text by section and page number for each newconcept that is introduced. Helpful Hints and Don't Try Its: Numeroususeful tips for solving problems, as well as warnings of commonstudent pitfalls. Unique Explanations: Conceptually clear, mathematicallyfairly simple, yet also sufficiently precise andrigorous. A more extensive set of reference materials, includingolder and newer editions of the major textbooks, as well as anumber of less commonly used titles, is available online at [ahref="http://www.conceptualthermo.com/](http://www.conceptualthermo.com/) <http://www.conceptualthermo.com/a>. Undergraduate and graduate students of chemistry, physics,engineering, geosciences and biological sciences will benefit fromthis book, as will students preparing for graduate school entranceexams and MCATs.

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Thermodynamics of Biochemical Reactions emphasizes the fundamental equations of thermodynamics and the application of these equations to systems of biochemical reactions. This emphasis leads to new thermodynamic potentials that provide criteria for spontaneous change and equilibrium under the conditions in a living cell.

This book is a physical chemistry textbook that presents theessentials of physical chemistry as a logical sequence from itsmost modest beginning to contemporary research topics. Many bookscurrently on the market focus on the problem sets with a cursorytreatment of the conceptual background and theoretical material,whereas this book is concerned only with the conceptual developmentof the subject. Comprised of 19 chapters, the book willaddress ideal gas laws, real gases, the thermodynamics of simplestems, thermochemistry, entropy and the second law, the Gibbsfree energy, equilibrium, statistical approaches to thermodynamics,the phase rule, chemical kinetics, liquids and solids, solutionchemistry, conductivity, electrochemical cells, atomic theory, wavemechanics of simple systems, molecular orbital theory, experimentaldetermination of molecular structure, and photochemistry and thetheory of chemical kinetics.

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